

## PRESS RELEASE

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Tokyo Institute of Technology research: Antiaromatic molecule displays record electrical conductance

(Tokyo, 19 July 2017) **Researchers demonstrate high electrical conductance for an antiaromatic nickel complex — an order of magnitude higher than for a similar aromatic complex. Since the conductance is also tunable by electrochemical gating, antiaromatic complexes are promising materials for future electronic devices.**

Organic materials often have a lower production cost than traditional electric conductors like metals and semiconductors. Not all organic systems conduct electricity well, however. A class of organic materials known as antiaromatic compounds — featuring planar rings of carbon atoms sharing a number of electrons that is a multiple of 4 — have been predicted to be excellent conductors, but this prediction has been hard to verify since antiaromatic molecules are usually unstable. Now, Shintaro Fujii and Manabu Kiguchi from the Tokyo Institute of Technology and colleagues have performed a systematic study of charge transport in a single, stable antiaromatic molecule. Compared to a structurally related aromatic molecule (where the carbon rings share an extra two electrons), its record electrical conductance is an order of magnitude higher.

The researchers studied a particular norcorrole-based nickel complex, Ni(nor), which is antiaromatic but stable, and a structurally similar aromatic, porphyrin-based nickel complex, Ni(porph). They measured the conductivities of the two compounds by means of the scanning tunneling microscopy break-junction technique; in such a setup, the current through a single molecule sandwiched between two parts of a broken junction is measured as a function of applied voltage. With a conductance of over  $4 \times 10^{-4}$  conductance quanta, Ni(nor) is the most conductive known organometallic complex. Ni(porph) was found to have a value about 25 times less, a result confirming the superior conductivity of antiaromatic molecules. Via theoretical calculations of the molecules' electronic structure and charge transport property, the scientists were able to identify the origin of the antiaromaticity-enhanced conductance: for Ni(nor), the lowest unoccupied molecular orbital lies closer to the Fermi level (the amount of work needed to add one electron to the system) than for Ni(porph).

Fujii and colleagues also succeeded in demonstrating tunability of the single-molecule conductance of Ni(nor). By applying a technique known as electrochemical gating, which enables controlling the molecular energy levels relative to the Fermi level of the source and drain electrodes (of the single-molecule junction) by varying an applied electrochemical potential, the researchers demonstrated a 5-fold modulation of the conductance of Ni(nor).

The findings of Fujii and colleagues show that antiaromatic materials are promising low-cost systems exhibiting high electrical conductance. In the words of the researchers, their study

“provides relevant guidelines for the design of molecular materials for highly conducting single-molecule electronics.

## Background

### Aromaticity and antiaromaticity (Figure 2)

Organic molecules are called aromatic when they feature a planar ring of carbon atoms with resonance bonds — a type of bond intermediate between a single and a double bond, resulting in high chemical stability (i.e., low reactivity). The archetypal aromatic molecule is benzene,  $C_6H_6$ , featuring a hexagonal aromatic ring of carbon atoms. The number of electrons (so-called  $\pi$ -electrons) shared by the ring is always a multiple of 4 plus 2 (for benzene, for example, it's 6), a property known as Hückel's rule.

Antiaromaticity is a similar property: a planar carbon ring, but with a number of  $\pi$ -electrons that is a multiple of 4 (for cyclobutadiene, for example, it's 4). Such a situation results in chemical instability.

Fujii and colleagues worked with a stable, nickel-based complex featuring an antiaromatic moiety. Compared to a similar complex with an aromatic moiety, its conductance is an order of magnitude larger, confirming the earlier prediction that antiaromatic molecules are excellent electric conductors.

### Conductance

How well a material conducts electricity is expressed via a quantity called electrical conductance,  $G$ . It is the inverse of electrical resistance,  $R = 1/G$ . Conductance is defined as the ratio of the current  $I$  running through the material (or, in the present work, a single molecule) and the voltage  $V$  across it. The unit of conductance is the siemens,  $S$ , but conductance values are often given with respect to the conductance quantum,  $G_0 \approx 7.7 \times 10^{-5} S$ .

Fujii and colleagues obtained values of the conductance for related antiaromatic and aromatic nickel-based complexes by measuring their current-voltage ( $I$ - $V$ ) characteristics in a scanning-tunneling microscopy (STM) setup.

### Reference

Shintaro Fujii<sup>1</sup>, Santiago Marqués-González<sup>1</sup>, Ji-Young Shin<sup>2</sup>, Hiroshi Shinokubo<sup>2</sup>, Takuya Masuda<sup>3</sup>, Tomoaki Nishino<sup>1</sup>, Narendra P. Arasu<sup>4</sup>, Héctor Vázquez<sup>4</sup>, Manabu Kiguchi<sup>1,\*</sup>, Highly-conducting molecular circuits based on antiaromaticity, *Nature Communications*, DOI: 10.1038/NCOMMS15984

1. Department of Chemistry, Graduate School of Science and Engineering, Tokyo Institute of Technology, Ookayama, Meguro-ku, Tokyo 152-8551, Japan
2. Department of Applied Chemistry, Graduate School of Engineering, Nagoya University, Aichi 464-8603, Japan
3. Global Research Center for Environment and Energy Based on Nanomaterials Science (GREEN), National Institute for Materials Science (NIMS), Tsukuba 305-0044, Japan
4. Institute of Physics, Academy of Sciences of the Czech Republic, Cukrovarnicka 10, Prague CZ-162 00, Czech Republic

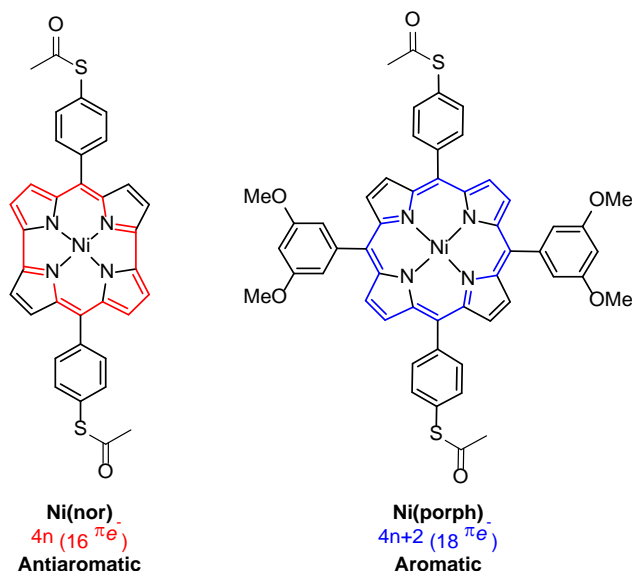
\*Corresponding author email: [fujii.s.af@m.titech.ac.jp](mailto:fujii.s.af@m.titech.ac.jp), [kiguti@chem.titech.ac.jp](mailto:kiguti@chem.titech.ac.jp)

## Contact

Emiko Kawaguchi  
Public relations section,  
Tokyo Institute of Technology  
E-mail: [media@jim.titech.ac.jp](mailto:media@jim.titech.ac.jp)  
Tel: +81-3-5734-2975

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**Figure 1.** Structures of the molecules used in the study of Fujii and colleagues.

left: Antiaromatic norcorrole-based Ni complex, Ni(nor). right: Aromatic Ni porphyrin-based complex, Ni(porph).

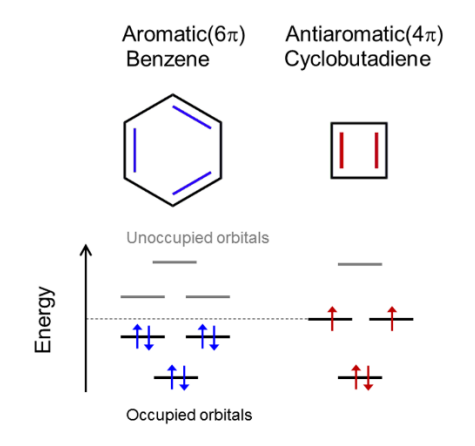


Figure 2. Molecular orbital energy levels for aromatic benzene and antiaromatic cyclobutadiene.